Balancing Net Ionic Equations

Write the TOTAL Net Ionic Equation for each of the following reactions:
Don't forget to balance and simplify them in the end!!!

1. $Br_{2\,(l)} + NaI_{(aq)} \rightarrow NaBr_{(aq)} + I_{2\,(s)}$

2. $Ca(OH)_{2\,(aq)} + HCl_{(aq)} \rightarrow CaCl_{2\,(aq)} + H_2O_{(l)}$

3. $Mg_{(s)} + AgNO_{3\,(aq)} \rightarrow Ag_{(s)} + Mg(NO_3)_{2\,(aq)}$

Use your solubility charts to identify which compound(s) are aqueous or not; and then balance and simplify the TOTAL Net Ionic Equation!

4. $AgNO_3 + KBr \rightarrow AgBr + KNO_3$

5. $Ni_{(s)} + Pb(NO_3)_2 \rightarrow Ni(NO_3)_2 + Pb_{(s)}$

6. $Ca + H_2O \rightarrow Ca(OH)_2 + H_2$

Predict the products for each of the following reactions, write-balance-simplify the Total Net Ionic Equation for each reaction. ALSO, identify the spectator ions in each reaction.

7. $AgNO_{3\,(aq)} + CaCl_{2\,(aq)} \rightarrow$

8. $Al_{(s)} + NiSO_{4\,(aq)} \rightarrow$

9. $AgNO_{3\,(aq)} + NaCl_{(aq)} \rightarrow$